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Chapter 9
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Section 2

Chapter 9:Section 2:
Ideal Stoichiometric
Calculations. Essential
questions. ...

Stoichiometry is the
part of chemistry that
applies the balanced
chemical equation to
determine the
quantities of reactants
and ... (O₂) have
different molar masses.
Mole ratios. A mole
ratio gives the
relationship between

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any two compounds in
a balanced chemical ...

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**Chapter 9: Section
2: Ideal
Stoichiometric
Calculations**

9-2 Ideal

Stoichiometric

Calculations Ideal

Stoichiometry - All

reactants are

converted into

products I. A Common

Method for Solving All

Stoichiometry

Problems A. Mass-Mass

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Problems 1. Start with a known mass of reactant or product, find an unknown mass of another reactant or product 2. All other stoichiometry problems are derivations ...

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SECTION 2. PROBLEMS

Write the answer on the line to the left.

Show all your work in the space provided. 1.

The following equation represents a laboratory preparation for oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow$

$2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many moles of O_2 form if 3.0 mol of KClO_3 are totally consumed? 2.

Given the following equation: $\text{H}_2(\text{g}) +$

$\text{F}_2(\text{g}) \rightarrow 2\text{HF}(\text{g})$

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3 PROBLEMS Write the

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the left. Show all your

work in the space

provided. 1. 88% The

actual yield of a

reaction is 22 g and

the theoretical yield is

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25 g. Calculate the percentage yield. 2.

6.0 mol of N_2 are mixed with 12.0 mol of H

mc06se cFMsr i-vi

SECTION 2 continued

Date Class _____ 60.2 9

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100. of lithium c a C ti.

l o c. i o g di l C10 c —

LCi(,; — h. The oxygen

gas produced in part

ahas density ot 1.43

g/L aiculate the olurne

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of this as.. 76

STOICHIOMETRY

MODERN CHEMISTRY a.

— . 81 g 6. A car air bag requires 70. L of nitrogen gas ...

Date. FCHAPJ

REVIEW.

Chapter 9 Intro to
Stoichiometry Section

9.1 . Chapter 9 9.1

Objectives • Define
stoichiometry. •

Describe the
importance of the mole
ratio in stoichiometric

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calculations. • Write a mole ratio relating two substances in a chemical equation. ...

Chapter 9 9.2

Objectives

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SECTION 1 Introduction to Stoichiometry

SECTION 2 Ideal Stoichiometric

Calculations SECTION 3

Limiting Reactants and Percentage Yield Why

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3 PROBLEMS Write the

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answer on the line to the left. Show all your work in the space provided. 1. 88% The

actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2.

6.0 mol of N_2 are mixed with 12.0 mol of H_2 according ...

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stoichiometry.

Stoichiometry is the branch of chemistry that deals with ; a) mass relationships of elements in compounds ; b) mass relationships between reactants and products in chemical reactions.

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2 Answers Chapter 9 -

Stoichiometry. 9-1

Introduction to
Stoichiometry.

Composition

Stoichiometry - deals
with mass relationships
of elements in
compounds Reaction

Stoichiometry -

Involves mass
relationships between
reactants and products
in a chemical reaction.

I. Reaction

Stoichiometry

Problems A, Four

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CHAPTER 9, Main Idea

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Ratios of substances in chemical reactions can be used as conversion factors. Key Terms
composition
stoichiometry

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