

Chapter 4 Atomic Structure

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90Chapter 4The Structure of the Atom Figure 4-5 Dalton's atomic theory explains the conservation of mass when a compound forms from its com- ponent elements. Atoms of ele- ments A and B combine in a simple whole-number ratio, in this case two B atoms for each A atom, to form a compound.

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On: Defining the Atom, Structure of the Nuclear Atom, Distinguishing Among Atoms. Excuse me while I add in my own comments.

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Chapter 4 - Atomic Structure. STUDY. PLAY ____ added to atomic theory through experiments with _____. Thomson, cathode ray tubes. Which model of the atom is thought to be true? Electrons are distributed around the nucleus and occupy almost all the volume of the atom.

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Section 4.2 Structure of the Nuclear Atom One change to Dalton's atomic theory is that atoms are divisible into subatomic particles: Electrons, protons, and neutrons are examples of these fundamental particles There are many other types of particles, but we will study these three

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Section 4.2 - Structure of an Atom. Protons, electrons, and neutrons are subatomic particles. A proton is a positively charged particle subatomic particle that is found in the nucleus of an atom. A proton has a +1 charge.

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Created by. frannieC97 GO. 4.1 Defining the Atom 4.2 Structure of the Nuclear Atom 4.3 Distinguishing between Atoms. Terms in this set (34) What is an atom? the smallest particle on an element that retains its identity in a chemical reaction.

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Chapter 4. Atomic Structure. Chapter 4 Pretest. 1. True or False: Compounds have fixed compositions. 2. What is an atom? 3. Which of the following units is a unit. of mass? a. mL b. °C. c. g d. cm. 4. What is density? 5. Which two of the following events can. take place when a liquid absorbs energy? a. The average kinetic energy of the

Chapter 4

Chapter 4 – Atomic Structure. Overview of the Atom and Atomic Structure. The Three Sub-Atomic Particles of the Atom. Proton +1. In the nucleus. Neutron. 0 (no charge) In the nucleus. Electron-1. In the electron cloud. Sub-Atomic Particle Name. Charge . Location in the atom. The nucleus.

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atomic mass unit: defined as the mass of a carbon -12 atom: nucleus: central core of an atom, which contains most of the atom's mass: group: a vertical column of elements in the periodic table: neutrons: subatomic particles with no charge: protons: positively charged subatomic particles: atom

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